



NOVEMBER EXAMINATION PAPER 2: CHEMISTRY

QUESTION 1

1.1 Which of the following is an example of a mixture?

- Element with the electron configurations: $1s^2, 2s^2, 2p^4$

- 1.1 The chemical formula of sodium sulfide

- | | |
|---|-------------------------|
| A | Fe_2S_3 |
| B | Fe_3S_4 |
| C | Fe_2S_5 |
| D | Fe_3S_2 |

- ☐ A Ion-Dipole
☐ B Dipole-dipole
☐ C Van der Waals
☐ D Hydrogen bond

- 1.5 The particles of a gas ...

- A are tightly packed.
- B vibrate against each other.
- C move freely and rarely touch.
- D move over one another.

- 1.6 Element with the highest electronegativity?

- ☐ A Hydrogen
☐ B Helium
☐ C Oxygen
☐ D Fluorine

- 1.7 Within an atom of Group 16, it rises to become atom 26

- A mass number of the atom will increase
B atomic number of the atom will decrease
C number of electrons will decrease
D number of electrons will increase

- 1.18 An element with the following Aufbau diagram will be classified as a



- A noble gas
B halogen
C alkali earth-metal
D alkali metal

Question Paper 2 November 2014 Physical Science

Jim Breithaupt



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