



Objectives

Class

Date

Study to provide an understanding of the environmental microbiology of the environment

Objectives

Practical 10: Microbiology of the Environment

Microbiology of the Environment

Microbiology of the Environment

Objectives

Practical 10: Microbiology of the Environment

Microbiology of the Environment

Microbiology of the Environment

1. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

2. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

3. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

4. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

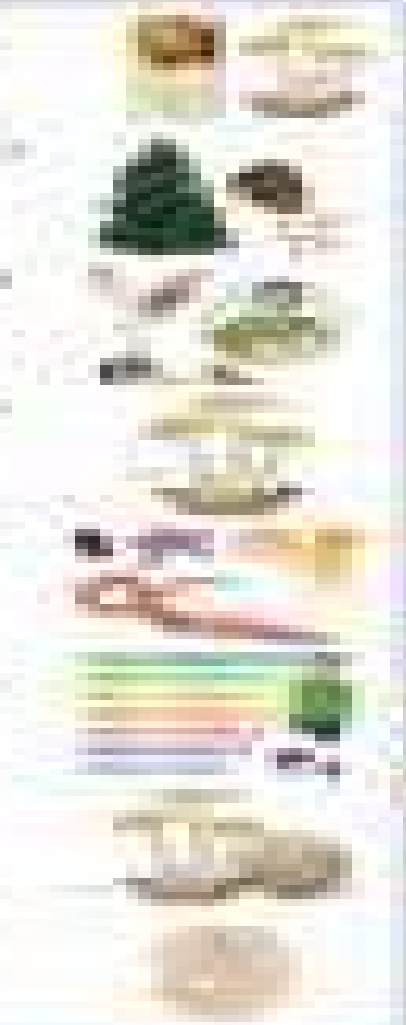
5. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

6. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

7. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

8. To understand the role of microorganisms in the environment and the importance of the environment in the life cycle of microorganisms.

Practical 10: Microbiology of the Environment



Amphibian Study Guide Answer Key

Christine Dugan



Amphibian Study Guide Answer Key:

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$$

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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c is the limiting reagent 4.33 g of h 2 are left over 13

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is
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which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate
what is the

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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