

Unit 7 Acids and Bases Study Guide

Chemistry Standard Set 5

Chapter 19

Key Vocabulary Terms

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|------------------------|-------------------------|----------------------|
| 1. Arrhenius Acid | 2. Arrhenius Base | 3. Hydrogen Donor |
| 4. Bronsted-Lowry Acid | 5. Bronsted-Lowry Base | 6. Hydrogen Acceptor |
| 7. Dissociation | 8. Partial Dissociation | 9. Hydronium |
| 10. Neutrality | 11. Buffer | 12. Amphoteric |
| 13. pH Scale | 14. Logarithmic Scale | |

Concepts

1. Observable properties of acids, bases, and salt solutions
2. Distinguish between the Arrhenius and the Bronsted-Lowry definitions of acid or bases.
3. An acid is a substance that, when dissolved in water, forms hydrogen ions (or protons, H^+). A base is a substance that, when dissolved in water, forms hydroxide ions (OH^-).
4. Distinguish between strong acids and bases that fully dissociate and weak acids and bases that partially dissociate.
5. Problems and sources of acid rain.
6. How antacids work in the body.

Items for Memorization

Litmus paper show that litmus is red in the presence of an acid and blue with a base. Phenolphthalein is colorless with an acid and pink with a base.

pH of pure water is 7 and amphoteric.

$pH + pOH = 14.000$; As the pH increases, the concentration of H^+ decreases.

Acids

- sour taste
- react with bases to form salt and water (neutralization)
- pH less than 7
- changes blue litmus paper to red
- changes bicarbonate indicator from red to yellow
- react with reactive metals (like iron, zinc, and aluminum) to produce hydrogen gas
- reacts with limestone ($CaCO_3$) to produce carbon dioxide.

Bases

- Taste bitter (like soap)
- feel slippery to the touch
- react with acids to form salt and water (neutralization)
- change red litmus paper to blue
- pH more than 7
- if soluble in water, known as alkali

Arrhenius concept. An acid is a substance that, when dissolved in water, forms hydrogen ions (or protons, H^+). A base is a substance that, when dissolved in water, forms hydroxide ions (OH^-). **Bronsted-Lowry concept.** An acid is a proton donor; a base is a proton acceptor.

Skills

1. Determine whether a substance is an acid or base based on its pH.
2. Identify acids and bases encountered in and around the home (baking soda, drain cleaner, vinegar, citrus fruits, battery acids, antacids, household ammonia)
3. Write simple equations to show how acids dissociate to release H^+ in solutions.
4. Ability to calculate pH, or pOH based on the hydrogen ion concentration.

Acids And Bases Study Guide Key

Calvin Yates

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Acids And Bases Study Guide Key:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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