

12) In the first half of the reaction, 0.010 mol of N_2O_4 is converted to NO_2 (total) with a volume of 0.250 L. What is the concentration?

$$\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$$

N_2O_4 is 0.010 mol
 NO_2 is 0.020 mol
 $0.020 \text{ mol} / 0.250 \text{ L} = 0.080 \text{ M}$

13) Predict whether each change will shift the equilibrium left, right, or no change will occur for the following equilibrium reaction: $\text{CO}_2(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{H}_2\text{CO}_3(aq) + \text{H}_2\text{O}(l)$

- a. ~~no change will occur~~
- b. ~~equilibrium shifts to the right~~
- c. ~~equilibrium shifts to the left~~
- d. ~~no change will occur~~
- e. ~~equilibrium shifts to the right~~

14) For each change, predict whether the reaction will shift left, right, or no change will occur.

a. ~~no change will occur~~

$\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$

b. ~~no change will occur~~

$\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$

15) Predict which species will be in solution after all precipitates are removed from the reaction. $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$

a. ~~Pb^{2+}~~

b. ~~Cl^{-}~~

c. ~~PbCl_2~~

d. ~~Pb^{2+} and Cl^{-}~~

$\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$
 $\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq) \rightleftharpoons \text{PbCl}_2(s)$

$\text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq)$

16) Predict all the following results which, along with their appropriate reason, result from the following reaction: $\text{Ba}^{2+}(aq) + \text{CO}_3^{2-}(aq) \rightleftharpoons \text{BaCO}_3(s)$

a. ~~no change will occur~~

b. ~~Ba^{2+} will increase~~

c. ~~BaCO_3 will increase~~

$\text{Ba}^{2+}(aq) + \text{CO}_3^{2-}(aq)$



$\text{Ba}^{2+}(aq) + \text{CO}_3^{2-}(aq)$

$\text{Ba}^{2+}(aq) + \text{CO}_3^{2-}(aq) \rightleftharpoons \text{BaCO}_3(s)$

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G Psacharopoulos



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