

MANUAL DE INSTRUCCIONES

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Acer Za3 Manual

Puja Mehta



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Libros españoles en venta ,1995
: *User's Guide* ,1991

Acer 913, User's Guide ,1988

ACER 1125E ,1990

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + 2\text{Cu} + 2\text{AgNO}_3 + 2\text{AgNO}_3$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$ $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 150.1 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{126.0 \text{ g}}{150.1 \text{ g}} \times 100 = 83.9\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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80 g i2o5 1 mol i2o5 1 mol i2 xs 1 333 8 g i2o5 1 mol i2o5 28 g co 1 mol co

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