

Note: Change in Page Numbering

Problem 21 (10 points)

Find the number of moles in 200. mL of 4.7 M HCl.

moles = 940

$$\begin{aligned} (M)(V) &= (4.7)(200) \\ &= 940 \text{ mol} \end{aligned}$$

Problem 22 (10 points)

Find the concentration of 50. mL of 12. M HCl when diluted to 200. mL.

M = 3

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2}$$

Problem 23 (10 points)

Find the Initial Volume necessary when 12. M HCl when diluted to make 400. mL

V = 40 ml of 12 M HCl

$$\frac{M_1 V_1 = M_2 V_2}{M_1}$$

$$V_1 = \frac{M_2 V_2}{M_1}$$

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