

Covalent Bonding Review Sheet – Dec 12, 2012

Name _____ class period _____

- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
- Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
- Since most compounds contain multiple bonds, a substance may contain more than one type of bond.

Section 1: Naming and Formula Writing

Write the formulas for the following covalent compounds

- | | |
|--------------------------|-----------------------------------|
| 1. Nitrogen Monoxide | <u>NO</u> |
| 2. Dinitrogen Monoxide | <u>N₂O</u> |
| 3. Nitrogen Dioxide | <u>NO₂</u> |
| 4. Dinitrogen Trioxide | <u>N₂O₃</u> |
| 5. Dinitrogen Tetraoxide | <u>N₂O₄</u> |
| 6. Dinitrogen Pentaoxide | <u>N₂O₅</u> |
| 7. carbon dioxide | <u>CO₂</u> |
| 8. phosphorus triiodide | <u>PI₃</u> |
| 9. sulfur dichloride | <u>SCl₂</u> |
| 10. nitrogen trifluoride | <u>NF₃</u> |

Write names from the following formulas

- | | |
|------------------------------------|------------------------------------|
| 11. N ₂ F ₄ | <u>dinitrogen tetrafluoride</u> |
| 12. SCl ₄ | <u>sulfur tetrachloride</u> |
| 13. ClF ₃ | <u>chlorine trifluoride</u> |
| 14. SiO ₂ | <u>silicon dioxide</u> |
| 15. P ₄ O ₁₀ | <u>tetraphosphorus decoxide</u> |
| 16. SiF ₄ | <u>silicon tetrafluoride</u> |
| 17. H ₂ O | <u>dihydrogen monoxide</u> |
| 18. Br ₂ | <u>dibromide</u> or <u>bromine</u> |
| 19. HBr | <u>hydrogen monobromide</u> |
| 20. N ₂ S ₃ | <u>dinitrogen trisulfide</u> |

Covalent Bonding Vocabulary Review Answer Key

BM King



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