

# GCSE Chemistry: Covalent Bonding Lesson & Activities

## Covalent Bonding

Specification Link:  
Chemical bonding: 4.1.1.4

### Task

Draw a diagram to show the electrons and the covalent bonds in the molecules below:

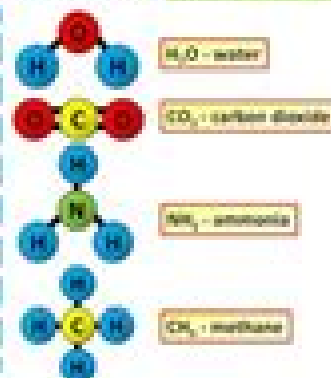
- |                                   |  |
|-----------------------------------|--|
| 1. $\text{CO}_2$ - carbon dioxide | 7. $\text{PCl}_5$ - phosphorus pentachloride |
| 2. $\text{CH}_4$ - methane        | 8. $\text{HCl}$ - hydrogen chloride          |
| 3. $\text{NH}_3$ - ammonia        | 9. $\text{SiF}_4$ - silicon tetrafluoride    |
| 4. $\text{H}_2\text{O}$ - water   | 10. $\text{NO}_2$ - nitrogen dioxide         |
| 5. $\text{O}_2$ - oxygen          | 11. $\text{SiO}_2$ - silicon dioxide         |
| 6. $\text{H}_2$ - hydrogen        | 12. $\text{CBr}_4$ - carbon tetrabromide     |

## Simple Covalent Compounds

Specification Link:  
Chemical bonding: 4.1.1.4

### Introduction

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. All have very strong bonds between the atoms, but much weaker forces holding the molecules together. When one of these substances melts or boils, it is these weak intermolecular forces that break, not the strong covalent bonds.



## Giant Covalent Compounds

Specification Link:  
Chemical bonding: 4.1.1.4

### Introduction

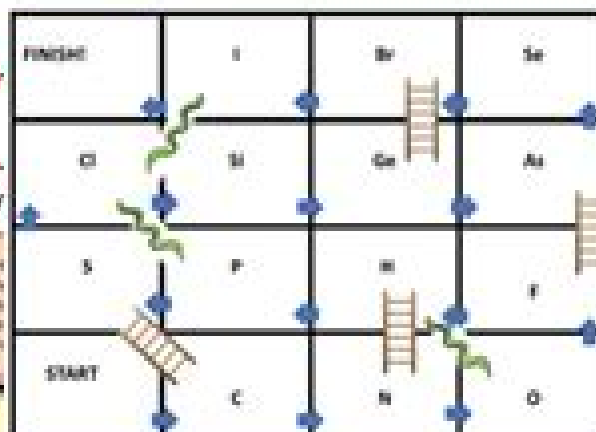
Giant covalent structures contain a lot of non-metal atoms, each joined to adjacent atoms by covalent bonds. The atoms are closely arranged into giant regular lattices - extremely strong structures because of the many bonds involved.

### Properties of giant covalent structures

- Very high melting points - Substances with giant covalent structures have very high melting points, because a lot of strong covalent bonds must be broken.
- Variable conductivity - Diamond does not conduct electricity. Graphite contains free electrons, so it does conduct electricity.



- Task**
1. Read the giant covalent structures.
  2. In pairs, draw a diagram of a giant covalent structure.
  3. Be ready to discuss with others.



## Simple molecular compounds SPAG activity

Covalent bonds form between non-metal atoms. Each bond consists of a shared pair of electrons and is very strong. Covalently bonded substances fall into two main types:

- Simple molecules
- Giant covalent structures

## Properties of simple molecular substances

**Low melting and boiling points** - This is because the weak intermolecular forces break down easily.

**Non-conductive** - Substances with a simple molecular structure do not conduct electricity. This is because they do not have any free electrons or an overall electric charge.

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. The strong covalent bonds between the atoms hold the molecules together. When these substances melt or boil, it is these weak 'intermolecular forces' that break, not the strong covalent bonds. Simple molecular substances are liquids or solids with low melting and boiling points.

Science

# Covalent Bonds Lesson

**Graeme Lofts**

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, positioned to the right of the author's name.

## Covalent Bonds Lesson :

*BIOLOGY | LIFE | Unit 1 | From Atoms to Cells* Smart Biology Inc.,2022-05-01 BIOLOGY LIFE UNIT 1 From Atoms to Cells focuses on the physical components that make up cells Embark on one continuous journey to understand and appreciate the interconnections between the subatomic atomic molecular macromolecular and cellular worlds We spend time covering the basics so you can understand the complex Moreover we explain the underlying why questions so you can truly understand This downloadable e book includes access to text over 350 high quality accurate figures 40 interactive structures and more to suit all learners Note Access to our online courseware including our animated video lessons is not included in this eBook but can be purchased at [www.smartbiology.com](http://www.smartbiology.com) E-chemistry Iii Tm (science and Technology)' 2003 Ed. ,

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