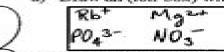
- An aqueous solution of rubidium phosphate (Rb₂PO₄) is mixed with an aqueous solution of magnesium nitrate and a precipitate forms.
 - a) Draw an (ion-box) with the four ions involved with this reaction. (1 mark)



b) Write a balanced formula equation for this reaction. Include all subscripts and coefficients. (2 marks)

2 Rb 3 PO+(+8) +3Mg(NO3)2(+9) - Mg (PO+)2(3) + 6 Rb NO3(-18)

 Write a balanced, complete (total) tonic equation for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)

6 Rb (og ,+ 2 PO + 3 Mg 2 tog + 6 NO 5 tog, - Mg (PO+) 2(5) + 6 Rb (og) + 6 NO 3 (og)

 Write a net ionic equation for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)

7

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- 1. Understanding the eBook Core Teaching Resources Chemistry Answer Key
 - The Rise of Digital Reading Core Teaching Resources Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Core Teaching Resources Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Core Teaching Resources Chemistry Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Core Teaching Resources Chemistry Answer Key

- Personalized Recommendations
- Core Teaching Resources Chemistry Answer Key User Reviews and Ratings
- Core Teaching Resources Chemistry Answer Key and Bestseller Lists
- 5. Accessing Core Teaching Resources Chemistry Answer Key Free and Paid eBooks
 - o Core Teaching Resources Chemistry Answer Key Public Domain eBooks
 - Core Teaching Resources Chemistry Answer Key eBook Subscription Services
 - Core Teaching Resources Chemistry Answer Key Budget-Friendly Options
- 6. Navigating Core Teaching Resources Chemistry Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Core Teaching Resources Chemistry Answer Key Compatibility with Devices
 - Core Teaching Resources Chemistry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Core Teaching Resources Chemistry Answer Key
 - Highlighting and Note-Taking Core Teaching Resources Chemistry Answer Key
 - Interactive Elements Core Teaching Resources Chemistry Answer Key
- 8. Staying Engaged with Core Teaching Resources Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Core Teaching Resources Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Core Teaching Resources Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Core Teaching Resources Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Core Teaching Resources Chemistry Answer Key
 - Setting Reading Goals Core Teaching Resources Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Core Teaching Resources Chemistry Answer Key

- Fact-Checking eBook Content of Core Teaching Resources Chemistry Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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