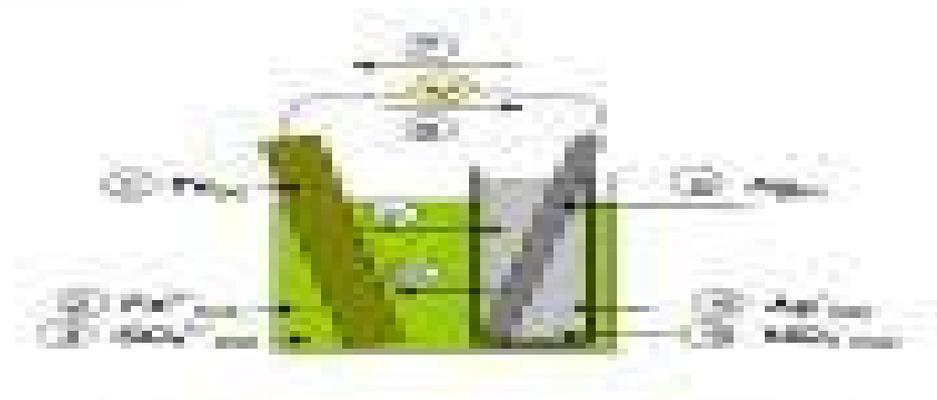


Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (2 marks)

- a. Oxidation: $Zn \rightarrow Zn^{2+} + 2e^{-}$
- b. Reduction: $Ag^{+} + e^{-} \rightarrow Ag$
- c. Anode: Zn
- d. Cathode: Ag
- e. Half-cell reactions: $Zn \rightarrow Zn^{2+} + 2e^{-}$ and $Ag^{+} + e^{-} \rightarrow Ag$
- f. Overall cell reaction: $Zn + 2Ag^{+} \rightarrow Zn^{2+} + 2Ag$

g. Write overall balanced equation:

$Zn + 2Ag^{+} \rightarrow Zn^{2+} + 2Ag$

h. Write overall cell equation: $Zn + 2AgNO_3 \rightarrow Zn(NO_3)_2 + 2Ag$

Source: <https://www.youtube.com/watch?v=8111111111>

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Donald W. Shive, Louise E. Shive



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