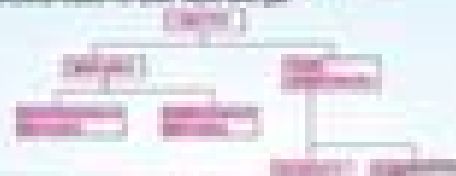


1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various sub-classifications of matter are given below:
- Matter is classified into pure and impure.
- It is classified into elements, compounds, mixtures, etc.
- It is classified into metals, non-metals, metalloids, etc.
- It is classified into solids, liquids, and gases.
- It is classified into pure and impure.
- It is classified into crystalline and amorphous.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are physical and chemical properties.
- Physical properties are those which can be measured without changing the chemical composition of the substance.
- Chemical properties are those which can be measured by changing the chemical composition of the substance.
- Colour, odour, melting point, density, etc. are physical properties.
- Combustion, corrosion, etc. are chemical properties.
- Quantitative measurement of physical properties is called measurement of matter.
- Measurement of matter is called measurement.
- Measurement of matter is called measurement.
- Measurement of matter is called measurement.

UNIT OF MEASUREMENT OF MATTER

Length (m)	Mass (kg)	Time (s)	Temperature (K)	Amount of substance (mol)
1000 m = 1 km	1000 g = 1 kg	1000 s = 1 min	1000 K = 1 °C	1000 mol = 1 kmol

UNIT OF MEASUREMENT OF MATTER

- 1. 1 m = 1000 mm, 1 km = 1000 m, 1 km = 1000 m, 1 km = 1000 m, 1 km = 1000 m.

3. SIGNIFICANT FIGURES

- Significant figures are those figures which are known with certainty and one which is uncertain or doubtful.
- Significant figures are those figures which are known with certainty and one which is uncertain or doubtful.

- Significant figures are those figures which are known with certainty and one which is uncertain or doubtful.

Rules for determining the number of significant figures

- 1. All non-zero digits are significant.
- 2. Zeros between two non-zero digits are significant.
- 3. Zeros at the end or right of a number are significant provided they are on the right side of a decimal point.
- 4. Leading zeros are not significant.
- 5. In a number written in scientific notation, all digits are significant.
- 6. Exact values are the values of certain measurements for which there is no uncertainty.
- 7. Rounding off the significant figures of a number to the first value of the next digit.
- 8. Addition and subtraction of significant figures: the result cannot have more digits to the right of the decimal point than either of the original numbers.
- 9. Multiplication and division of significant figures: the result must be rounded off to the same number of significant figures as the least number of significant figures in the original numbers.

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