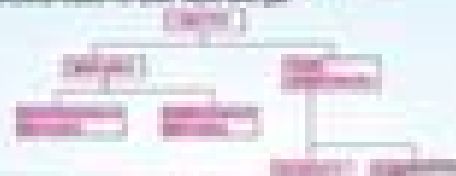


1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various properties of matter are used to differentiate between the two.
- Matter is classified into pure and impure.
- Homogeneous mixture: composition is uniform throughout.
- Heterogeneous mixture: composition is not uniform.
- Colloid: particles of substance are dispersed in another substance.
- Suspension: particles of substance are dispersed in another substance.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are physical and chemical properties.
 - Physical properties are those which can be measured without changing the identity of the substance.
 - Chemical properties are those which can be measured by changing the identity of the substance.
 - Colour, odour, melting point, density, etc. are some physical properties.
 - Combustion, corrosion, etc. are some chemical properties.
 - Quantitative measurement of physical properties is necessary for scientific study.
 - Properties of matter are measured in SI units.
 - Some important quantities and their units are:
- | Quantity | SI Unit |
|---------------------------|-------------------------------|
| Mass (g) | Kilogram (kg) |
| Volume (cm ³) | Cubic metre (m ³) |
| Temperature (°C) | Kelvin (K) |

UNIT CONVERSIONS

1 kg = 1000 g	1 g = 10 ⁻³ kg	1 m = 100 cm	1 cm = 10 ⁻² m	1 L = 1000 mL
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UNIT CONVERSIONS

- 1 kg = 1000 g, 1 g = 1000 mg, 1 mg = 10⁻³ g, 1 L = 1000 mL, 1 mL = 10⁻³ L, 1 m = 100 cm, 1 cm = 10⁻² m, 1 dm = 10⁻¹ m, 1 nm = 10⁻⁹ m, 1 μm = 10⁻⁶ m, 1 Å = 10⁻¹⁰ m, 1 pm = 10⁻¹² m, 1 fm = 10⁻¹⁵ m, 1 attom = 10⁻¹⁸ m, 1 zepto = 10⁻²¹ m, 1 yocto = 10⁻²⁴ m, 1 ronto = 10⁻²⁷ m, 1 quecto = 10⁻³⁰ m, 1 septo = 10⁻³³ m, 1 octo = 10⁻³⁶ m, 1 nono = 10⁻³⁹ m, 1 deca = 10, 1 hecto = 100, 1 kilo = 1000, 1 mega = 10⁶, 1 giga = 10⁹, 1 tera = 10¹², 1 peta = 10¹⁵, 1 exa = 10¹⁸, 1 zetta = 10²¹, 1 yotta = 10²⁴.

3. SIGNIFICANT FIGURES

- Significant figures are those digits in a number which are known with certainty and one which is uncertain or doubtful.

- Significant figures are those digits which are known with certainty and one which is uncertain or doubtful.

Rules for determining the number of significant figures:

- Non-zero digits are significant.
- Zeros between two non-zero digits are significant.
- Zeros at the end or right of a number are significant provided they are on the right side of a decimal point.
- Leading zeros are not significant figures.
- In a number written in scientific notation, all digits are significant.
- Exact values are the numbers of atoms, molecules, etc. for a substance.
- Round-off to the nearest integer of a number is the true value of the result.
- In addition and subtraction of significant figures, the result cannot have more digits to the right of the decimal point than either of the original numbers.
- In multiplication and division of significant figures, the result must be rounded off to the same number of significant figures as the least number of significant figures in the original numbers.

Aakash Chemistry Lab Manual

Yan Bai



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1. Understanding the eBook Aakash Chemistry Lab Manual
 - The Rise of Digital Reading Aakash Chemistry Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Aakash Chemistry Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aakash Chemistry Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aakash Chemistry Lab Manual
 - Personalized Recommendations
 - Aakash Chemistry Lab Manual User Reviews and Ratings
 - Aakash Chemistry Lab Manual and Bestseller Lists
5. Accessing Aakash Chemistry Lab Manual Free and Paid eBooks
 - Aakash Chemistry Lab Manual Public Domain eBooks
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 - Aakash Chemistry Lab Manual Compatibility with Devices
 - Aakash Chemistry Lab Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aakash Chemistry Lab Manual
 - Highlighting and Note-Taking Aakash Chemistry Lab Manual
 - Interactive Elements Aakash Chemistry Lab Manual
8. Staying Engaged with Aakash Chemistry Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aakash Chemistry Lab Manual
9. Balancing eBooks and Physical Books Aakash Chemistry Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aakash Chemistry Lab Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Aakash Chemistry Lab Manual
 - Setting Reading Goals Aakash Chemistry Lab Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Aakash Chemistry Lab Manual
 - Fact-Checking eBook Content of Aakash Chemistry Lab Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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