

of substance:

Substance:

of particles of a substance
in mol

6.02×10^{23} particles
amount of substance

Constant:

particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

Unit:

mol, measured in g
equivalent to mol

Atomic Structure

Isotopes:

→ Same # of protons, same element
of p^+ and e^- but different
neutrons different mass

Physical Properties:

→ Can be different - e.g. different
melting + boiling points
→ i.e. heavy water has higher melting
point + boiling point + density

Chemical Properties:

→ Same no. of electrons + same electronic
structure in outer shell + transfer of e^- is
unaffected hence have same chemical
properties.
→ Different no. of n has no effect
on bonding + reactions

Compounds + Ionic Formulas:

Alkali NO_3^-

Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure
- Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$

At RTP (20°C and 100kPa):
(293K and 100000Pa)

is $24\text{ dm}^3 \text{ mol}^{-1}$
at RTP.

Volume /
 dm^3

Electronegativity:

→ the measure of attraction of a bonded atom
in a covalent bond

Electronegativity increases across period:

- Increase in no. of protons in an atom
→ Nuclear charge increases

Electronegativity DE:

- More electron shells
→ More shielding
→ Atomic radius increases
→ Nuclear charge decreases
→ Nuclear attraction decreases
→ Electronegativity decreases

OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS



→ Consists of a small dense
nucleus of protons + neutrons
surrounded by electrons orbiting
in electron shells in shells

- No. of p^+ = No. of e^-

→ Equal + opposite charges cancel
in atom has neutral charge
- Neutrons important in nuclear
stability - Electrostatic repulsion
between p^+ in small dense nucleus

Mass Number:

→ No. of p^+ + n in nucleus of atom
- p^+ and n have relative mass of 1
and e^- mass is negligible is 0

Relative Isotopic Mass:

→ The mass of an atom of an isotope
relative to $1/12$ th the mass of an atom of ^{12}C

Relative Atomic Mass:

→ The weighted mean mass of an atom of
an element relative to $1/12$ th the mass of
an atom of ^{12}C

Weighted Mean Mass:

→ Average mass of all isotopes, taking into
account the relative abundance of each isotope

- It dipoles
Molecule is
out + inlets

e.g. H_2O

- O is more electronegative than H
- O-H has permanent dipole since do not cancel there is an over.

e.g. CO_2

O is more electronegative than C
 $\text{C}=\text{O}$ has a dipole
Molecule has no net dipole

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Richard Parsons



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Advanced Economics Through Diagrams Andrew Gillespie, 2001 DT These highly successful revision guides have been brought right up to date for the new A Level specifications introduced in September 2000 DT Oxford Revision Guides are highly effective for both individual revision and classroom summary work The unique visual format makes the key concepts and processes and the links between them easier to memorize DT Students will save valuable revision time by using these notes instead of condensing their own DT In fact many students are choosing to buy their own copies so that they can colour code or highlight them as they might do with their own revision notes *My Revision Notes: OCR A Level Chemistry A* Mike Smith, 2017-02-06 Exam Board OCR Level A Level Subject Chemistry First Teaching September 2015 First Exam Summer 2016 With *My Revision Notes OCR A Level Chemistry A* you can Manage your own revision with step by step support from experienced teacher and examiner Mike Smith Apply biological terms accurately with the help of definitions and key words Plan and pace your revision with the revision planner Test understanding with questions throughout the book Get exam ready with last minute quick quizzes available on the Hodder Education website *OCR(B) AS/A2 Chemistry (Salters) Student Unit Guide: Units F333 and F336 Chemistry in Practice and Individual Investigation* Frank Harriss, 2011-12-09 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements This guide offers advice on preparing for the OCR B Salters Units F333 Chemistry in Practice and F336 Individual Investigation assessments The Practical Chemistry section deals with experimental results and practical techniques It explains how to obtain and record results and how to process them [OCR AS/A Level Year 1 Chemistry A Student Guide: Modules 1 and 2](#) Mike Smith, 2015-10-09 Exam Board OCR Level AS A level Subject Chemistry First Teaching September 2015 First Exam Summer 2016 Written by experienced author Mike Smith this Student Guide for Chemistry Helps you identify what you need to know with a concise summary of the topics examined in the AS and A level specifications Consolidates understanding with tips and knowledge check questions Provides opportunities to improve exam technique with sample answers to exam style questions Develops independent learning and research skills Provides the content for generating individual revision notes

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Table of Contents A2 Chemistry Ocr Rates Practical Exam Quantitative

1. Understanding the eBook A2 Chemistry Ocr Rates Practical Exam Quantitative
 - The Rise of Digital Reading A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Advantages of eBooks Over Traditional Books
2. Identifying A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an A2 Chemistry Ocr Rates Practical Exam Quantitative
 - User-Friendly Interface
4. Exploring eBook Recommendations from A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Personalized Recommendations

- A2 Chemistry Ocr Rates Practical Exam Quantitative User Reviews and Ratings
- A2 Chemistry Ocr Rates Practical Exam Quantitative and Bestseller Lists
- 5. Accessing A2 Chemistry Ocr Rates Practical Exam Quantitative Free and Paid eBooks
 - A2 Chemistry Ocr Rates Practical Exam Quantitative Public Domain eBooks
 - A2 Chemistry Ocr Rates Practical Exam Quantitative eBook Subscription Services
 - A2 Chemistry Ocr Rates Practical Exam Quantitative Budget-Friendly Options
- 6. Navigating A2 Chemistry Ocr Rates Practical Exam Quantitative eBook Formats
 - ePub, PDF, MOBI, and More
 - A2 Chemistry Ocr Rates Practical Exam Quantitative Compatibility with Devices
 - A2 Chemistry Ocr Rates Practical Exam Quantitative Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Highlighting and Note-Taking A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Interactive Elements A2 Chemistry Ocr Rates Practical Exam Quantitative
- 8. Staying Engaged with A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers A2 Chemistry Ocr Rates Practical Exam Quantitative
- 9. Balancing eBooks and Physical Books A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection A2 Chemistry Ocr Rates Practical Exam Quantitative
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Setting Reading Goals A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of A2 Chemistry Ocr Rates Practical Exam Quantitative
 - Fact-Checking eBook Content of A2 Chemistry Ocr Rates Practical Exam Quantitative

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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composition ... Core logging: Optimizing best practice (Part One). We must not forget that geotechnical core logging
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