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oxide and nitric acid magnesium oxide nitric acid magnesium nitrate

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type of chemical reaction for this reaction we have a neutralization reaction balancing strategies here we have a
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web the balanced chemical equation for this reaction is $\text{mg s} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg no}_3 \text{ 2 aq} + \text{h}_2 \text{ g}$ in this equation the s symbolizes
solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the
gaseous hydrogen produced

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equation balancer balanced chemical equation $2\text{hno}_3 \text{ mg c}_2\text{h}_3\text{o}_2 \text{ 2 2hc}_2\text{h}_3\text{o}_2 \text{ mg no}_3 \text{ 2}$

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reaction where one mole of solid magnesium mg and two moles of aqueous nitric acid hno 3 react to form one mole of
aqueous magnesium nitrate mg no 3 2 and one mole of dihydrogen h 2 gas show chemical

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molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of
nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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