

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $2\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Phillips, Cheryl Wistrom, 2000 Chemical Misconceptions Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources *Fundamentals of the Optics of Materials* Vladimir I. Gavrilenko, Volodymyr S. Ovechko, 2023-10-19 This textbook showcases the rapidly developing field of materials optics It is aimed at a broad audience as the readers require only university entry level knowledge of physics chemistry and optics It overviews the basics of optical engineering and the typical and widely used applications of materials optics with the first general chapters corresponding to the standard university courses targeting bachelor s and master s degrees in physics The next few chapters present the modern developments in materials optics such as nano plasmonics nano photonics and optical properties of nano sized materials intended for readers familiar with the basic elements of quantum mechanics Some more specialized chapters address recent developments in fields such as optics of solid surfaces plasma optics optics of composites alloys and metamaterials optics of anisotropic materials optics of organic and biological materials and relativistic effects in optics The appendices present a more advanced description of selected topics with important reference materials subject index and extended list of publications as well as numerous examples and problems to better orient readers interested in gaining further knowledge of the subject **Exceptional Child Education Resources** , 2001

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