



CHEMISTRY SSC-I

Revised Syllabus

Time allowed: 2:40 Hours

Total Marks Sections B and C: 53

NOTE:- Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separately provided answer book. Use supplementary answer sheet i.e. Sheet-B if required. Write your answers neatly and legibly.

SECTION – B (Marks 33)

Q. 2 Answer any ELEVEN parts. The answer to each part should not exceed three to four lines. (11 x 3 = 33)

- (i) Identify the branch of Chemistry, that is related to the following information:
 - a. Sulphuric Acid is weaker than HCl.
 - b. Gases can be compressed by applying pressure.
 - c. Hair contain a special class proteins called keratins which is present in nails and wool.
- (ii) Determine the formula masses of the following:
 - a. Urea $(\text{NH}_2)_2\text{CO}$
 - b. Ammonium Nitrate NH_4NO_3
- (iii) What were the conclusions of Rutherford's experiment?
- (iv) Write any three uses of Isotopes?
- (v) Identify the **group** and **period** of $^{27}_{13}\text{Al}$, ^9_5B , $^{24}_{12}\text{Mg}$ on the basis of electronic configuration.
- (vi) What is Atomic size? Write its trends in periodic table.
- (vii) What is the importance of Glue and Adhesives in our society?
- (viii) Differentiate between Diffusion and Effusion of Gases. Also give examples.
- (ix) Why does evaporation lower the temperature of Liquids?
- (x) Define **Saturated**, **Unsaturated** and **Super-saturated** solutions. Also give examples.
- (xi) Differentiate between Colloid and Suspension, with the help of examples.
- (xii) Define Oxidation Reduction reaction. Also give one example from daily life.
- (xiii) Draw a labelled diagram of the Daniel's cell.
- (xiv)
 - a. Which of the given elements is more metallic and Why? Mg or Al.
 - b. What is the importance of Mg?
 - c. F_2 oxidizes all the halide ions to free Halogens. How?
- (xv) Compare reactivity of Alkali and Alkaline earth metals by their positions in the periodic table.

SECTION – C (Marks 20)

Note: Attempt any TWO questions. All questions carry equal marks.

(2 x 10 = 20)

- | | | |
|-------------|--|------|
| Q. 3 | a. Define Electro-negativity. Explain the Electronegative characters of elements. Give examples. | (04) |
| | b. Give any three important uses of Platinum. | (03) |
| | c. What is Alloying? Also give example. | (02) |
| | d. Define Rusting of Iron. | (01) |
| Q. 4 | a. How can you manufacture Sodium Hydroxide from Brine? | (04) |
| | b. A solution is prepared by dissolving 10 gram of Haemoglobin in water to make up 1dm^3 in volume. Calculate molarity of this solution. Molar mass of Haemoglobin is $6.51 \times 10^4 \text{ g / mole}$. | (04) |
| | c. What is Concentration Unit? Give examples. | (02) |
| Q. 5 | a. What is Vapour Pressure? Explain effects of temperature on Vapour Pressure. | (04) |
| | b. Define Intermolecular forces. Explain Hydrogen Bonding found in Water. Also draw diagram. | (04) |
| | c. What is Shielding effect? Which atom has greater shielding effect and Why? Be or Mg. | (02) |

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Sebastian Brünink



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mothers and develop treatment plans that address their unique risks and challenges included in this article are some of the more common nanda plans for postpartum care students general students care plan

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