

A blank periodic table grid with the following structure:

- s-block:** The first two columns (Groups 1 and 2).
- p-block:** The last six columns (Groups 13 to 18).
- d-block:** The ten columns between the s-block and p-block (Groups 3 to 10).
- f-block:** Two rows of seven columns each, located below the main body of the table, corresponding to the lanthanide and actinide series.

The grid is labeled with 's-block', 'p-block', 'd-block', and 'f-block' in their respective regions. The rows are numbered 1 through 7 on the left side.

1. Obtain a card with electron configurations for your group as assigned by your teacher.
 - a. Record your team number _____ and team name _____.
2. Locate where your set of elements should be in Model 1.
 - a. Write the last orbital notation in the electron configuration for each element in your set in its respective box.
 - b. What is the relationship between your answer in Question 2a and the "block" of the table where your set of elements is located?
3. What is the relationship between the last orbital notation in your set of ground state electron configurations and the rows (horizontal) numbered on the left-hand side of the periodic table in Model 1?

Cracking The Periodic Table Code Worksheet Answers

Theodore S. Abbgy



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