

CHEMICAL BONDS

IONIC, COVALENT, & METALLIC

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READING COMPREHENSION & QUESTIONS

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Types of Chemical Bonds Reading Comprehension

What is the difference between a molecule and a compound?

What of the substances shown in this diagram are ionic?

What of the substances shown in this diagram are covalent?

Chemical Bonds and Electronegativity

Electronegativity is an atom's ability to attract electrons. Elements with low electronegativity have outer shells that are almost empty and tend to gain or lose electrons easily. Elements with high electronegativity have outer shells that are almost full and tend to hold on to their electrons. Ionic bonds form when the electronegativity of two atoms differs significantly, and covalent bonds form when electronegativity of the atoms are more similar. Typically, the greater the difference in electronegativity between two elements, the more ionic the bond. The greater the similarity in electronegativity, the more covalent the bond.

Electronegativity increases from bottom left to top right on the periodic table.

Which atom has a higher electronegativity?

Which set is a cation and which set is an anion?

Which set of compounds are ionic?

Which set of compounds are covalent?

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