

Read chapter 8, 2.5-2.7

- Explain the difference between metallic, ionic, and covalent bonding.  
Metallic- cations share a sea of electrons.  
Ionic- atoms give and take electrons.  
Covalent- atoms share electrons.
- Why can metals conduct electricity?  
electrons are free to move around
- Predict the chemical formula of the ionic compound formed between the following pairs of elements.
 

a. Al and F	b. K and S	c. Mg and N	d. Ba and O
$\text{AlF}_3$	$\text{K}_2\text{S}$	$\text{Mg}_3\text{N}_2$	$\text{BaO}$
- Define the term lattice energy and tell what factors govern the magnitude of the lattice energy of an ionic compound. The amount of energy that is required to break an ionic bond. The magnitude depends on the charge of the ion and the size of the ion. smaller ions and more charge have higher lattice energy.
- Arrange the following substances according to their expected lattice energies, listed them from lowest lattice energy to highest: LiCl, KCl, KBr, CaO  
 $\text{KBr} < \text{KCl} < \text{LiCl} < \text{CaO}$
- The lattice energies of NaF and MgO are given in table 8.2. Account for the difference in these two quantities.  
MgO has higher charges so the lattice energy is higher
  - Account for the difference in the lattice energies of  $\text{MgCl}_2$  and  $\text{SrCl}_2$ , which are also listed in the table.  
Mg is smaller than Sr, so the lattice is greater
- Explain the following trends in lattice energy
 

a. $\text{MgO} > \text{MgS}$ Oxygen is smaller than S	b. $\text{LiF} > \text{CsBr}$ Cs is larger than Li Br is larger than F	c. $\text{CaO} > \text{KF}$ CaO has larger charges than KF
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- How is bonding in  $\text{Cl}_2$  different than NaCl?  
In  $\text{Cl}_2$ , the electrons are being shared between the two chlorine atoms, but in NaCl, the Cl is taking the electron from the sodium
- Explain why a Br atom in  $\text{Br}_2$  satisfies the octet rule whereas a free Br atom does not.  
Br has 7 valence electrons, but when it bonds to itself they both contribute an electron to share so they each feel 8 electrons around them.
- Draw the Lewis structure for  $\text{O}_2$ . The bond in  $\text{O}_2$  is shorter than the O-O single bond. Explain this observation.  
 $\text{O}=\text{O}$   
A double bond is shorter than a single bond
- Predict whether the following compounds are molecular or ionic
 

a. $\text{B}_2\text{H}_6$	b. $\text{CH}_3\text{OH}$	c. $\text{LiNO}_3$	d. $\text{Sc}_2\text{O}_3$	e. CsBr	f. NOCl	g. $\text{Ag}_2\text{SO}_4$
Molecular	molecular	ionic	ionic	ionic	covalent	ionic

# Covalent Bond Chapter 8 Worksheet 1 Answer Key

**ML Morrison**



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