

Chemistry Chapter 8

Bonding and Molecular Structure

Key idea: Arrangement of nuclei and electrons of the bonded atoms results in a lower energy than for the separate atoms

- **Coulomb's Law:**
 - Coulomb's law says that the attractive force between ions of opposite charge is directly proportional to the product of the ion charges and inversely proportional to the square of the distance between them
 - $F = (q_A \times q_B) / r^2$
 - q_A and q_B = ion charges
 - R = distance between ions.
 - Negative number = attractive force, so disregard negative sign and biggest number = bigger magnitude.
- **Lewis Structures:**
 - **Key to bonding:** electron sharing to achieve a full valence shell
 - **Octet rule:** electrons are distributed in such a way that each element is surrounded by 8 electrons, an octet
 - Exceptions:
 - For H_2 molecule stability is reached with 2 electrons.
 - For expanded octets/valence: elements have more than 8 electrons (often 10 or 12) in a Lewis structure.
 - 3rd row and below can do this.
 - Incomplete Octet: elements have less than 8 electrons in a Lewis structure.
 - Molecules containing H, B and Be.
 - Non-bonding valence electrons (usually in pairs) are called lone pairs
 - Each dot represents a valence electron
 - The element symbol represents the nucleus and core electrons.
 - Hydrogen can only make one bond.
 - For Oxygen containing acids, the Hydrogen of the acid is attached to the Oxygen.
 - For organic molecules:
 - **Carbon forms 4 bonds** (a double bond is 2 bonds) and is generally in the interior of the molecule.
 - **Hydrogen forms only 1 bond** and is therefore a terminal atom (on the outside of the molecule).
 - **Oxygen usually forms 2 bonds.**
 - **Nitrogen usually forms 3 bonds** (may form 4 in a cation like NH_4^+).
- **Electronegativity (χ)**
 - The ability of an atom in a molecule to attract (pull) electrons to itself. The less electronegativity, the more bonds it forms and will be the central atom in a complex molecule.
- **Formal Charge (FC)**

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