

Tutorial 1: Chemical Kinetics Answers

1. Consider the following concentration-time data for the reaction of iodide ion and hypochlorite ion (OCl^-). The products are chloride ion and hypiodite ion (OI^-).

Experiment	Time (s)	$[\text{I}^-]$	$[\text{OCl}^-]$	$[\text{OH}^-]$
1	0	2.40×10^{-4}	1.60×10^{-3}	1.00
	10	2.16×10^{-4}	1.36×10^{-3}	1.00
2	0	1.20×10^{-4}	1.60×10^{-3}	1.00
	10	1.08×10^{-4}	1.48×10^{-3}	1.00
3	0	2.40×10^{-4}	4.00×10^{-5}	1.00
	10	2.34×10^{-4}	3.40×10^{-5}	1.00
4	0	1.20×10^{-4}	1.60×10^{-3}	2.00
	10	1.14×10^{-4}	1.54×10^{-3}	2.00

- (a) Write a balanced equation for the reaction.
 (b) Determine the rate law, and calculate the value of the rate constant.
 (c) Does the reaction occur by a single-step mechanism? Explain.
 (d) Propose a mechanism that is consistent with the rate law, and express the rate constant in terms of the rate constants for elementary steps in your mechanism. (Hint: Transfer of an H^+ ion between H_2O and OCl^- is a rapid reversible reaction.)



- b) assume that $\text{rate} = \frac{-\Delta[\text{I}^-]}{\Delta t} = k[\text{OCl}^-]^a[\text{I}^-]^b[\text{OH}^-]^c$
 compare experiments 1 and 3:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{2.4 \times 10^{-4} - 2.34 \times 10^{-4}} = \left(\frac{1.6 \times 10^{-3}}{4 \times 10^{-5}} \right)^a$$

$$4^a \approx 4$$

$$\therefore a = 1$$

compare experiments 1 and 2:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{1.2 \times 10^{-4} - 1.08 \times 10^{-4}} = \left(\frac{2.4}{1.2} \right)^b$$

$$2^b \approx 2$$

$$\therefore b = 1$$

compare experiments 1 and 4:

$$\frac{2.4 \times 10^{-4} - 2.17 \times 10^{-4}}{1.2 \times 10^{-4} - 1.14 \times 10^{-4}} = 2 \left(\frac{1}{2} \right)^c$$

$$2 \left(\frac{1}{2} \right)^c \approx 4$$

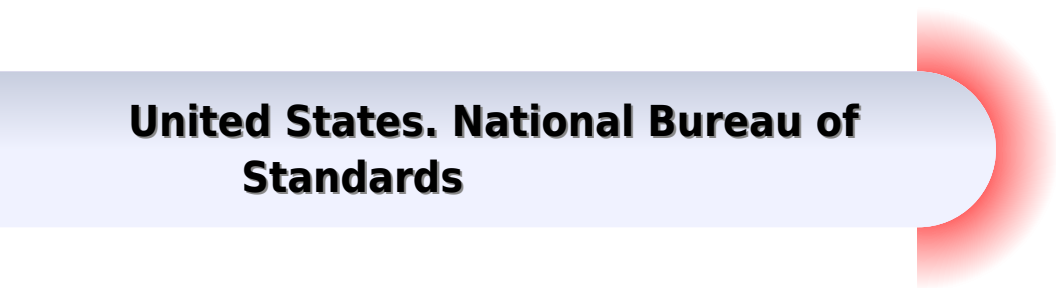
$$\therefore c = -1$$

$$\text{rate} = \frac{-\Delta[\text{I}^-]}{\Delta t} = k \frac{[\text{OCl}^-][\text{I}^-]}{[\text{OH}^-]}$$

$$\frac{1.2 \times 10^{-4} - 1.08 \times 10^{-4}}{10} = k \frac{1.2 \times 10^{-4}(1.6 \times 10^{-3})}{1}$$

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Table of Contents 22 1 Review And Reinforcement Chemical Kinetics Answers

1. Understanding the eBook 22 1 Review And Reinforcement Chemical Kinetics Answers
 - The Rise of Digital Reading 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 22 1 Review And Reinforcement Chemical Kinetics Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 22 1 Review And Reinforcement Chemical Kinetics Answers

- Personalized Recommendations
 - 22 1 Review And Reinforcement Chemical Kinetics Answers User Reviews and Ratings
 - 22 1 Review And Reinforcement Chemical Kinetics Answers and Bestseller Lists
5. Accessing 22 1 Review And Reinforcement Chemical Kinetics Answers Free and Paid eBooks
 - 22 1 Review And Reinforcement Chemical Kinetics Answers Public Domain eBooks
 - 22 1 Review And Reinforcement Chemical Kinetics Answers eBook Subscription Services
 - 22 1 Review And Reinforcement Chemical Kinetics Answers Budget-Friendly Options
 6. Navigating 22 1 Review And Reinforcement Chemical Kinetics Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 22 1 Review And Reinforcement Chemical Kinetics Answers Compatibility with Devices
 - 22 1 Review And Reinforcement Chemical Kinetics Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Highlighting and Note-Taking 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Interactive Elements 22 1 Review And Reinforcement Chemical Kinetics Answers
 8. Staying Engaged with 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 22 1 Review And Reinforcement Chemical Kinetics Answers
 9. Balancing eBooks and Physical Books 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 22 1 Review And Reinforcement Chemical Kinetics Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Setting Reading Goals 22 1 Review And Reinforcement Chemical Kinetics Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 22 1 Review And Reinforcement Chemical Kinetics Answers

- Fact-Checking eBook Content of 22 1 Review And Reinforcement Chemical Kinetics Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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