

Ⓢ Diese Waffe ist frei erwerbbar ab 18 Jahre oder älter. Zu Ihrer eigenen und zur Sicherheit anderer lesen Sie bitte die Bedienungsanleitung vollständig und sorgfältig durch, bevor Sie mit der Waffe den ersten Versuch zu schießen unternehmen. Käufer und Besitzer haben die Pflicht, alle Regeln über den Gebrauch und den Besitz von Luftpistolen zu befolgen.

Ⓢ You must be 18 years or older to purchase this CO₂ gun. Carefully read and understand these instructions before using this weapon. Buyer and user are obliged to follow the instructions for the safe use of airguns.

Ⓢ Cette arme se peut acquérir à partir de l'âge de 18 ans. Afin de garantir votre propre sécurité et la sécurité d'autrui veuillez lire complètement et soigneusement le mode d'emploi, avant d'entreprendre le premier essai de tirer avec l'arme. Les acheteurs et possesseurs sont obligés de respecter toutes les règles concernant l'emploi et la possession de pistolets à air comprimé.

Ⓢ Esta arma se puede adquirir a partir de los 18 años. Antes de hacer el primer ensayo de tiro sírvase leer completamente y cuidadosamente las instrucciones de empleo para su seguridad propia y para la seguridad de otros. Compradores y poseedores están obligados a observar todas las reglas respecto al uso y a la posesión de pistolas de aire.

Gebrauchsanleitung - Instruction manual - Mode d'emploi - Instrucciones para el uso



CO₂-Luftpistole

CO₂ pistol

Pistolet à air comprimé de CO₂

Pistola de aire de gas CO₂

Mod. PPK/S

cal. 4.5mm / .177 BB

(Stahlrundkugeln)

(steel BB's)

(balles rondes d'acier)

(balas redondas de acero)



Walther Ppk S Co2 User Manual

N Noddings



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instead of ml we have $v_m = \frac{32 \text{ g mol}^{-1}}{1790 \text{ g l}^{-1}} = 0.0405 \text{ l mol}^{-1}$ nonnumber

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web this volume is known as the molar gas volume v_m the volume of a gas may be calculated from its number of moles using volume of gas moles v_m rearranging this moles of gas $\frac{v}{v_m}$

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web jun 8 2023 $v = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{K} = 22.4 \text{ l}$ in other words 1 mole of a gas will occupy 22 4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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