Hess's Law Worksheet

 Calculate ΔH for the reaction: C₂H₄ (g) + H₂ (g)→C₂H₅ (g), from the following data.

 $\Delta H = -1411. \text{ kJ}$

$$C_2H_6(g) + 3\% O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l)$$

 $\Delta H = -1560$. kJ

$$H_2(g) + \% O_2(g) \rightarrow H_2O(l)$$

 $\Delta H = -285.8 \text{ kJ}$

 Calculate ΔH for the reaction 4 NH₃ (g) + 5 O₂ (g) EB4 NO (g) + 6 H₂O (g), from the following data.

$$N_2(g) + O_2(g) \rightarrow 2 NO(g)$$

 $\Delta H = -180.5 \text{ kJ}$

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$

 $\Delta H = -483.6 \text{ kJ}$

Find ΔH for the reaction 2H₂(g) + 2C(s) + O₂(g) → C₂H₅OH(I), using the following thermochemical data.

 $\Delta H = -875. \text{ kJ}$

$$C(s) + O_2(g) \rightarrow CO_2(g)$$

 $\Delta H = -394.51 \, kJ$

 $\Delta H = -285.8 \text{ kJ}$

Calculate
☐H for the reaction CH₄ (g) + NH₃ (g) → HCN (g) + 3 H₂ (g), given:

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$

 $\Delta H = -74.9 \text{ kJ}$

$$H_2(g) + 2C(s) + N_2(g) \rightarrow 2HCN(g)$$

 $\Delta H = +270.3 \text{ kJ}$

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