

GCSE Chemistry: Covalent Bonding Lesson & Activities

Covalent Bonding

Specification Link:
Chemical bonding: 4.1.1.4

Task

Draw a diagram to show the electrons and the covalent bonds in the molecules below:

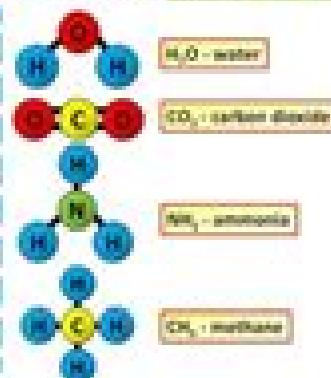
- | | |
|-----------------------------------|--|
| 1. CO_2 - carbon dioxide | 7. PCl_5 - phosphorus pentachloride |
| 2. CH_4 - methane | 8. HCl - hydrogen chloride |
| 3. NH_3 - ammonia | 9. SiF_4 - silicon tetrafluoride |
| 4. H_2O - water | 10. NO_2 - nitrogen dioxide |
| 5. O_2 - oxygen | 11. SnCl_4 - tin(IV) chloride |
| 6. H_2 - hydrogen | 12. CBr_4 - carbon tetrabromide |

Simple Covalent Compounds

Specification Link:
Chemical bonding: 4.1.1.4

Introduction

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. All have very strong bonds between the atoms, but much weaker forces holding the molecules together. When one of these substances melts or boils, it is these weak intermolecular forces that break, not the strong covalent bonds.



Giant Covalent Compounds

Specification Link:
Chemical bonding: 4.1.1.4

Introduction

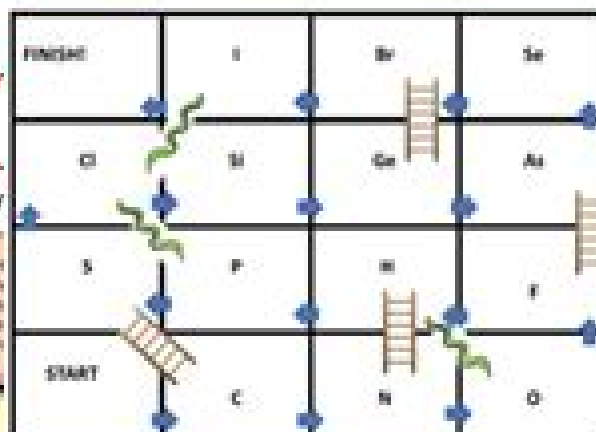
Giant covalent structures contain a lot of non-metal atoms, each joined to adjacent atoms by covalent bonds. The atoms are closely arranged into giant regular lattices - extremely strong structures because of the many bonds involved.

Properties of giant covalent structures

- Very high melting points - Substances with giant covalent structures have very high melting points, because a lot of strong covalent bonds must be broken.
- Variable conductivity - Diamond does not conduct electricity. Graphite contains free electrons, so it does conduct electricity.



- Task**
1. Read the giant covalent structures.
 2. In pairs, draw a diagram of a giant covalent structure.
 3. Be ready to discuss with others.



Simple molecular compounds SPAG activity

Covalent bonds form between non-metal atoms. Each bond consists of a shared pair of electrons and is very strong. Covalently bonded substances fall into two main types:

- Simple molecules
- Giant covalent structures

Properties of simple molecular substances

Low melting and boiling points - This is because the weak intermolecular forces break down easily.

Non-conductive - Substances with a simple molecular structure do not conduct electricity. This is because they do not have any free electrons or an overall electric charge.

Hydrogen, ammonia, methane and water are all simple molecules with covalent bonds. The strong covalent bonds between the atoms hold the molecules together. When these substances melt or boil, it is these weak 'intermolecular forces' that break, not the strong covalent bonds. Simple molecular substances are usually liquids or solids with low melting and boiling points.

Science

Covalent Bonds Lesson

Kaplan Test Prep



Covalent Bonds Lesson :

BIOLOGY | LIFE | Unit 1 | From Atoms to Cells Smart Biology Inc.,2022-05-01 BIOLOGY LIFE UNIT 1 From Atoms to Cells focuses on the physical components that make up cells Embark on one continuous journey to understand and appreciate the interconnections between the subatomic atomic molecular macromolecular and cellular worlds We spend time covering the basics so you can understand the complex Moreover we explain the underlying why questions so you can truly understand This downloadable e book includes access to text over 350 high quality accurate figures 40 interactive structures and more to suit all learners Note Access to our online courseware including our animated video lessons is not included in this eBook but can be purchased at [www smart biology com](http://www.smartbiology.com) **E-chemistry Iii Tm (science and Technology)' 2003 Ed.**

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Homework Helpers: Chemistry Greg Curran,2025-09-12 Homework Helpers Chemistry is a user friendly review book that will make every student or parent trying to help their child feel like he or she has a private Chemistry tutor Concepts are explained in clear easy to understand language and problems are worked out with step by step methods that are easy to follow Each lesson comes with numerous review questions and answer keynotes that explain each correct answer and why it s correct This book covers all of the topics in a typical one year Chemistry curriculum including A systematic approach to

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Overcoming Students' Misconceptions in Science Mageswary Karpudewan,Ahmad Nurulazam Md Zain,A.L. Chandrasegaran,2017-02-28 This book discusses the importance of identifying and addressing misconceptions for the successful teaching and learning of science across all levels of science education from elementary school to high school It suggests teaching approaches based on research data to address students common misconceptions Detailed descriptions of how these instructional approaches can be incorporated into teaching and learning science are also included The science education literature extensively documents the findings of studies about students misconceptions or alternative conceptions about various science concepts Furthermore some of the studies involve systematic approaches to not only creating but also implementing instructional programs to reduce the incidence of these misconceptions among high school science students These studies however are largely unavailable to classroom practitioners partly because they are usually found in various science education journals that teachers have no time to refer to or are not readily available to them In response this book offers an essential and easily accessible guide

[Roadmap to the Regents](#) Sasha Alcott,2003 If Students Need to Know It It s in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test s structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry

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Curriculum, 3e learnON and Print Pascale Warnant, 2025-10-20 **College Chemistry I** Norman H. Nachtrieb, 1960

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