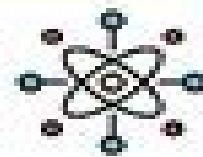


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the attraction between positive metal ions and a sea of delocalized electrons. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding occurs when two atoms share one or more pairs of electrons. This results in a stable molecule. Ionic bonding occurs when one atom transfers an electron to another atom, creating a positive ion and a negative ion. These ions are attracted to each other. Metallic bonding occurs in metals, where the atoms are arranged in a regular lattice, and the outer electrons are shared among all the atoms, forming a 'sea of electrons'.

Science Real

A simplified periodic table of elements, color-coded by groups. It shows the arrangement of elements from Hydrogen to Oganesson.

IONIC BONDING

Atoms combine to form ionic compounds by transferring electrons. This process is called bonding. The resulting ionic compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
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A diagram showing the formation of an ionic bond. It depicts a sodium atom (Na) with one valence electron and a chlorine atom (Cl) with seven valence electrons. An arrow shows the transfer of the sodium electron to the chlorine atom, resulting in a sodium ion (Na+) and a chloride ion (Cl-). The ions are then shown attracted to each other by electrostatic forces.

GUIDED READING ACTIVITY

Covalent Bonds Reading Notetaking Guide

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1. Understanding the eBook Covalent Bonds Reading Notetaking Guide
 - The Rise of Digital Reading Covalent Bonds Reading Notetaking Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonds Reading Notetaking Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonds Reading Notetaking Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonds Reading Notetaking Guide
 - Personalized Recommendations
 - Covalent Bonds Reading Notetaking Guide User Reviews and Ratings

- Covalent Bonds Reading Notetaking Guide and Bestseller Lists
- 5. Accessing Covalent Bonds Reading Notetaking Guide Free and Paid eBooks
 - Covalent Bonds Reading Notetaking Guide Public Domain eBooks
 - Covalent Bonds Reading Notetaking Guide eBook Subscription Services
 - Covalent Bonds Reading Notetaking Guide Budget-Friendly Options
- 6. Navigating Covalent Bonds Reading Notetaking Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonds Reading Notetaking Guide Compatibility with Devices
 - Covalent Bonds Reading Notetaking Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonds Reading Notetaking Guide
 - Highlighting and Note-Taking Covalent Bonds Reading Notetaking Guide
 - Interactive Elements Covalent Bonds Reading Notetaking Guide
- 8. Staying Engaged with Covalent Bonds Reading Notetaking Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonds Reading Notetaking Guide
- 9. Balancing eBooks and Physical Books Covalent Bonds Reading Notetaking Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonds Reading Notetaking Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonds Reading Notetaking Guide
 - Setting Reading Goals Covalent Bonds Reading Notetaking Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonds Reading Notetaking Guide
 - Fact-Checking eBook Content of Covalent Bonds Reading Notetaking Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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