

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                    |                                      |
|------------------------------------|--------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pOH} = -\log[\text{OH}^-]$ |
| 3. $\text{pH} + \text{pOH} = 14$   | 4. $\text{pH} + \text{pOH} = 14$     |

## Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the problem. Express answers in scientific notation when appropriate.

1. What substance has  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter hydrogen ion?
2. Autoionization constant  $K_w = 1.0 \times 10^{-14}$  at  $25^\circ\text{C}$ . What is the concentration of  $\text{OH}^-$  in a solution with  $[\text{H}^+] = 1.0 \times 10^{-3}$  moles/liter?
3. Autoionization constant  $K_w = 1.0 \times 10^{-14}$  at  $25^\circ\text{C}$ . What is the concentration of  $\text{H}^+$  in a solution with  $[\text{OH}^-] = 1.0 \times 10^{-3}$  moles/liter?
4. Autoionization constant  $K_w = 1.0 \times 10^{-14}$  at  $25^\circ\text{C}$ . What is the concentration of  $\text{H}^+$  in a solution with  $[\text{OH}^-] = 1.0 \times 10^{-10}$  moles/liter?
5. Autoionization constant  $K_w = 1.0 \times 10^{-14}$  at  $25^\circ\text{C}$ . What is the concentration of  $\text{H}^+$  in a solution with  $[\text{OH}^-] = 1.0 \times 10^{-12}$  moles/liter?
6. Autoionization constant  $K_w = 1.0 \times 10^{-14}$  at  $25^\circ\text{C}$ . What is the concentration of  $\text{H}^+$  in a solution with  $[\text{OH}^-] = 1.0 \times 10^{-14}$  moles/liter?

7.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
8.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
9.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
10.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
11.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
12.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
13.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
14.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
15.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
16.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
17.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter
18.  $[\text{H}^+] = 1.0 \times 10^{-7}$  moles/liter

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Huangqi Zhang**



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