

- *This is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.*
- *DO NOT USE A CALCULATOR. You may use ONLY the green periodic table.*
- *Try to work at a pace of 1.2 min per question. Time yourself. It is important that you practice working for speed.*
- *Then when time is up, continue working and finish as necessary.*

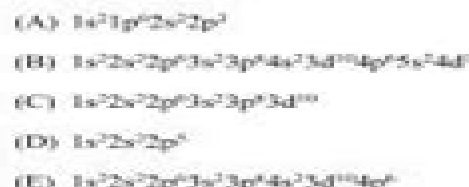
- Which of the following contains only atoms that are diamagnetic in their ground state?
 - Kr, Cu, and P
 - Cl, Mg, and Cd
 - Ar, K, and Ba
 - He, Sr, and C
 - Ni, Be, and Zn
- A valence electron from an arsenic atom might have an electron with the following set of quantum numbers in the ground state.
 - $n=4$; $L=1$; $m_l=0$; $m_s=+\frac{1}{2}$
 - $n=4$; $L=1$; $m_l=2$; $m_s=-\frac{1}{2}$
 - $n=3$; $L=1$; $m_l=0$; $m_s=+\frac{1}{2}$
 - $n=5$; $L=1$; $m_l=-1$; $m_s=-\frac{1}{2}$
 - $n=4$; $L=2$; $m_l=+1$; $m_s=+\frac{1}{2}$
- Calcium reacts with element X to form an ionic compound. If the state electron configuration of X is $1s^2 2s^2 2p^4$, what is the simplest formula for this compound?
 - CaX
 - CaX_2
 - Ca_2X_3
 - Ca_3X_2
 - Ca_2X_3
- The ground-state configuration of Fe^{2+} is which of the following?
 - $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$
 - $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$
 - $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
 - $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
 - $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
- In the ground state the highest-energy electron of a rubidium atom might have which of the following sets of quantum numbers?
 - $n=5$; $L=0$; $m_l=1$; $m_s=+\frac{1}{2}$
 - $n=5$; $L=1$; $m_l=1$; $m_s=+\frac{1}{2}$
 - $n=4$; $L=0$; $m_l=0$; $m_s=+\frac{1}{2}$
 - $n=5$; $L=0$; $m_l=0$; $m_s=+\frac{1}{2}$
 - $n=6$; $L=0$; $m_l=0$; $m_s=+\frac{1}{2}$

The next four questions refer to the following orbital diagrams.



- The least reactive element is represented by:
- The transition metal is represented by:
- The most chemically reactive element is represented by:
- The element in an excited state is represented by:

Use the following ground-state electron configurations for the following four questions.



- The electron configuration of a halogen is:
- This is a possible configuration for a transition metal atom.
- This electron configuration is not possible.
- This is a possible configuration of a transition metal ion.

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