

# Boyles law questions

## SCH3U Boyle's Law Worksheet

1. State the pressure-volume law both in words and in the form of an equation.  
2. To compress nitrogen at 1 atm from 750 mL to 500 mL, what must the new pressure be if the temperature is kept constant? **1.5 atm**
3. If oxygen at 128 kPa is allowed to expand at constant temperature until its pressure is 101.3 kPa, how much larger will the volume become? **1.26**
4. A sample of nitrogen at 101.3 kPa with a volume of 100 mL is carefully compressed at constant temperature in successive changes in pressure, equalling 5 kPa at a time, until the final pressure is 133.3 kPa. Calculate each new volume and prepare a plot of P versus V, showing P on the horizontal axis.
5. A sample of nitrogen at 20°C was compressed from 300 mL to 0.360 mL and its new pressure was found to be 400.0 Pa. What was the original pressure in kPa?  **$4.8 \times 10^{-4}$**
6. The pressure on 6.0 L of a gas is 200 kPa. What will be the volume if the pressure is doubled, keeping the temperature constant? **3L**
7. What would be the new volume if the pressure on 600 mL is increased from 90 kPa to 150 kPa?  
**360 mL**
8. A student collects 25 mL of gas at 96 kPa. What volume would this gas occupy at 101.325 kPa. There is no change in temperature or mass. **23.68 mL**
9. A gas measuring 525 mL is collected at 104.66 kPa. What volume does this gas occupy at 99.33 kPa? **553.17 mL**
10. A mass of gas occupies 1 L at 1 atm. At what pressure does this gas occupy  
a) 2 litres, **0.5 atm**  
b) 0.5 litres? **2 atm**
11. From the data in the following table calculate the missing quantity (assuming constant temperature).  
a)  $V_1 = 22.4 \text{ L}$ ;  $P_1 = 1 \text{ atm}$ ;  $P_2 = ? \text{ atm}$ ;  $V_2 = 2.8 \text{ L}$  **8 atm**  
b)  $V_1 = 60 \text{ mL}$ ;  $P_1 = ? \text{ kPa}$ ;  $P_2 = 101.3 \text{ kPa}$ ;  $V_2 = 16 \text{ mL}$  **27 kPa**  
c)  $V_1 = ? \text{ m}^3$ ;  $P_1 = 40 \text{ Pa}$ ;  $P_2 = 100 \text{ kPa}$ ;  $V_2 = 1.0 \text{ L}$   **$2.5 \times 10^9 \text{ m}^3$**   
d)  $V_1 = 2.50 \text{ L}$ ;  $P_1 = 7.5 \text{ atm}$ ;  $P_2 = ? \text{ atm}$ ;  $V_2 = 100 \text{ mL}$  **187.5 atm**

# Boyles Law Cavalcade Publishing Answers

**L Reisser**



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## **Table of Contents Boyles Law Cavalcade Publishing Answers**

1. Understanding the eBook Boyles Law Cavalcade Publishing Answers
  - The Rise of Digital Reading Boyles Law Cavalcade Publishing Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Boyles Law Cavalcade Publishing Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Boyles Law Cavalcade Publishing Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Boyles Law Cavalcade Publishing Answers
  - Personalized Recommendations
  - Boyles Law Cavalcade Publishing Answers User Reviews and Ratings
  - Boyles Law Cavalcade Publishing Answers and Bestseller Lists
5. Accessing Boyles Law Cavalcade Publishing Answers Free and Paid eBooks

- Boyles Law Cavalcade Publishing Answers Public Domain eBooks
- Boyles Law Cavalcade Publishing Answers eBook Subscription Services
- Boyles Law Cavalcade Publishing Answers Budget-Friendly Options
- 6. Navigating Boyles Law Cavalcade Publishing Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Boyles Law Cavalcade Publishing Answers Compatibility with Devices
  - Boyles Law Cavalcade Publishing Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Boyles Law Cavalcade Publishing Answers
  - Highlighting and Note-Taking Boyles Law Cavalcade Publishing Answers
  - Interactive Elements Boyles Law Cavalcade Publishing Answers
- 8. Staying Engaged with Boyles Law Cavalcade Publishing Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Boyles Law Cavalcade Publishing Answers
- 9. Balancing eBooks and Physical Books Boyles Law Cavalcade Publishing Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Boyles Law Cavalcade Publishing Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Boyles Law Cavalcade Publishing Answers
  - Setting Reading Goals Boyles Law Cavalcade Publishing Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Boyles Law Cavalcade Publishing Answers
  - Fact-Checking eBook Content of Boyles Law Cavalcade Publishing Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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