

Bond Energy

Chem Worksheet 16-2

Name _____

Bond energy is defined as the amount of energy required to break a bond. These values are positive, indicating that bond breaking is endothermic. Bond energies are reported in kilojoules per mole (kJ/mol). The energy for breaking a hydrogen-hydrogen bond is 436 kJ/mol so when a hydrogen-hydrogen bond is formed the process releases 436 kJ/mol.

In a chemical reaction several bonds are broken and formed. For example in the reaction below a hydrogen-hydrogen bond is broken and a fluorine-fluorine bond is broken. Two hydrogen-fluorine bonds are formed. The overall energy change for this process is calculated below.

example



$$\begin{aligned}\Delta H_{\text{rxn}} &= [\text{energy used for breaking bonds}] - [\text{energy formed in making bonds}] \\ &= [436 \text{ kJ/mol} + 155 \text{ kJ/mol}] - [2(567 \text{ kJ/mol})] \\ &= -543 \text{ kJ/mol}\end{aligned}$$



Bond energy: 436 kJ/mol



Bond Energy: 155 kJ/mol



Bond Energy: 567 kJ/mol

Average Bond Energies (kJ/mol)

| | | | | | |
|------|------------|-------|------------|-----|-------------|
| H-H | 436 kJ/mol | C-H | 413 kJ/mol | C=C | 614 kJ/mol |
| H-Cl | 431 kJ/mol | C-C | 348 kJ/mol | C≡C | 839 kJ/mol |
| H-F | 567 kJ/mol | C-N | 293 kJ/mol | C=O | 799 kJ/mol |
| N-H | 391 kJ/mol | C-O | 358 kJ/mol | O=O | 495 kJ/mol |
| N-O | 201 kJ/mol | C-F | 485 kJ/mol | C=O | 1072 kJ/mol |
| O-H | 463 kJ/mol | C-Cl | 328 kJ/mol | C=N | 615 kJ/mol |
| O-O | 146 kJ/mol | C-S | 259 kJ/mol | N=N | 418 kJ/mol |
| F-F | 155 kJ/mol | Cl-Cl | 242 kJ/mol | N≡N | 941 kJ/mol |
| | | | | C≡N | 891 kJ/mol |

Estimate the enthalpy change (ΔH_{rxn}) of the following reactions using the bond energies above.



Draw Lewis structures for reactants and products. Estimate the enthalpy change (ΔH_{rxn}) for the reactions using bond energies.



Bond Energy Extension Questions Answers

Rob Ritchie, Dave Gent



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