

Lewis base: a substance that donates electrons to form a covalent bond

Distinguishing Acids from bases

- Acids increase the concentration of H^+ in aqueous solution
- bases decreases the concentration of H^+ in aqueous solution

The Special role of the Hydronium ion

Hydronium ion

- Formula: H_3O^+ (aq)
- Lewis structure

Ions form: The Arrhenius Concept

Arrhenius definitions

- **Arrhenius acid** a substance that when dissolved in water, increases the concentration of H_3O^+ or (H^+)
- **Arrhenius base** a substance that when dissolved in water increases the concentration of OH^-

What the substance does: The Bronsted Lowry Concept

Bronsted Lowry Definitions used on proton transfer reactions

- **Bronsted Lowry acid**
 - A proton donor
 - $HCl \rightarrow Cl^- + H^+$
- **Bronsted Lowry base**
 - A proton acceptor
 - $NH_3 + H^+ \rightarrow NH_4^+$

Conjugate Acid base Pairs

Activity 7 5 Bronsted Lowry Acids And Bases Answers

Léo Pomar



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