

17A. A Solubility Product Constant

Introduction

The interaction of a slightly soluble ionic compound with its dissolved ions leads to another type of equilibrium (Ebbing/Gammon, Chapter 17). The equilibrium constant for this kind of equilibrium has a special name. It is called the *solubility product constant*.

Purpose

In this experiment, you will determine the solubility of $\text{Ca}(\text{IO}_3)_2$ (calcium iodate) and calculate its solubility product constant. You will also demonstrate that the solubility of $\text{Ca}(\text{IO}_3)_2$ is independent of the amount of water that is used to obtain a saturated solution of this substance.

Concept of the Experiment

If the solubility of a slightly soluble ionic compound is measured, the solubility product constant for the substance can be calculated (Ebbing/Gammon, Example 17.3). This is the method that you will adopt for $\text{Ca}(\text{IO}_3)_2$ in this experiment. The equilibrium between solid $\text{Ca}(\text{IO}_3)_2$ and its ions in a saturated solution is



If some analytical technique is used to determine the concentration of either Ca^{2+} ions or IO_3^- ions in the saturated solution, the solubility of $\text{Ca}(\text{IO}_3)_2$ will be known, and the solubility product constant (K_{sp}) can be calculated.

In this experiment, you will determine the concentration of IO_3^- ions through a titration with a standardized solution of $\text{Na}_2\text{S}_2\text{O}_3$ (sodium thiosulfate) in the presence of KI (potassium iodide), using starch as an indicator. A definition of standardization can be found in the experiment "An Acid-Base Titration Curve." Potassium iodide will react with the iodate ions to give I_2 as the sole product containing iodine. The reaction is



The molecular iodine reacts with $\text{S}_2\text{O}_3^{2-}$ ions during the titration according to



where $\text{S}_4\text{O}_6^{2-}$ is the tetrathionate ion. When these two equations are combined, multiplying the second reaction by 3 so the intermediate I_2 is eliminated, the net reaction occurring in the titration is obtained.



Starch is used as an indicator in this titration because it reacts with I_2 reversibly to form a dark blue color. I_2 is consumed in the titration, so the color fades as the titration progresses. You will know when a stoichiometric volume of the $\text{Na}_2\text{S}_2\text{O}_3$ solution has been added because at that point, one drop of that solution will cause the disappearance of the last trace of the blue color. A trial titration will enable you to find the approximate volume that is required before you do the first of two exact titrations.

A Solubility Product Constant Lab 17a Answers

JA Banks



A Solubility Product Constant Lab 17a Answers:

This book delves into A Solubility Product Constant Lab 17a Answers. A Solubility Product Constant Lab 17a Answers is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into A Solubility Product Constant Lab 17a Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
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 - Chapter 3: A Solubility Product Constant Lab 17a Answers in Everyday Life
 - Chapter 4: A Solubility Product Constant Lab 17a Answers in Specific Contexts
 - Chapter 5: Conclusion
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