

Atomic Structure – Ch. 3

PART A – SUBATOMIC PARTICLES

The table below contains information about several elements. In each case, enough information has been provided for you to fill in the blanks. Assume all atoms are neutral.

Isotope Name	Nuclear Symbol	Atomic Number	Mass Number	# of Protons	# of Electrons	# of Neutrons
1. calcium-40						
2.		12	24			
3.				1		2
4.	$^{197}_{79}\text{Au}$					
5.					26	30
6.			201	80		
7.		17				18

PART B – AVERAGE ATOMIC MASS

- Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.
- Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.
- Distinguish between mass number, relative atomic mass, and average atomic mass.

Atomic Structure Periodic Table Answers

Chris Sunley, Sam Goodman



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Chemistry: Concepts and Problems Clifford C. Houk, Richard Post, 1996-03-09 CHEMISTRY SECOND EDITION The fast easy way to master the fundamentals of chemistry Have you ever wondered about the differences between liquids gases and solids Or what actually happens when something burns What exactly is a solution An acid A base This is chemistry the composition and structure of substances composing all matter and how they can be transformed Whether you are studying chemistry for the first time on your own want to refresh your memory for a test or need a little help for a course this concise interactive guide gives you a fresh approach to this fascinating subject This fully up to date edition of Chemistry Concepts and Problems Has been tested rewritten and retested to ensure that you can teach yourself all about chemistry Requires no prerequisites Lets you work at your own pace with a helpful question and answer format Lists objectives for each chapter you can skip ahead or find extra help if you need it Reinforces what you learn with chapter self tests

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